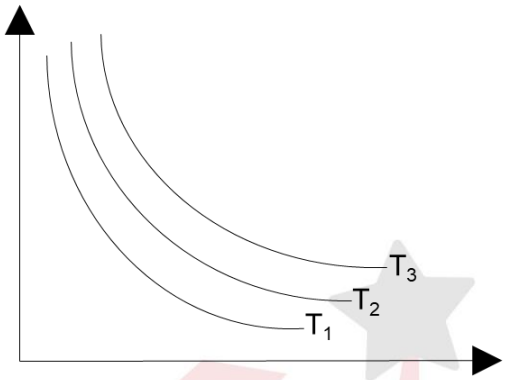
(ECC02-0030M-PMC-08H)

- "P" versus "V" curves and constant temperatures T1, T2, & T3 are given in figure. Which is correct?



A) $T_1 < T_2 < T_3$

B) $T_1 = T_2 = T_3$

C) $T_2 < T_1 < T_3$

D) $T_1 > T_2 > T_3$

(MCC02-UHS-01)

- Under what conditions of temperature and pressure will a real gas behave like an ideal gas?

Options	Temperature	Pressure
A)	Low	Low
B)	Low	High
C)	High	High
D)	High	Low

(MCC02-UHS-03)

- Which one of the following gases shows easily deviate from ideal behavior?

A) O_2

B) CO_2

C) N_2

D) H_2

(MCC02-UHS-04)

- Under high pressure which of the following gas show more ideality in character
  - A) N_2
  - B) He
  - C) NH_3
  - D) SO_2

(MCC02-UHS-05)

- Which of the statement is applicable for both ideal and real gases molecules?
  - A) Have no forces of attraction
  - B) Collisions between the molecules is elastic
  - C) Molecules are in random movement
  - D) The actual volume of gas is negligible as compared to the volume of gas

(MCC02-UHS-06)

- Gases deviate from ideal behavior at high pressure. Which of the following is correct for non-ideality?
  - A) At high pressure, the gas molecules move in one direction only
  - B) At high pressure, the collision between the gas molecules is increased manifold
  - C) At high pressure, the volume of the gas becomes insignificant
  - D) At high pressure, the inter-molecular attractions become significant

(MCC02-UHS-07)

- The molecules of gas show more deviation from ideal behavior at low temperature, because
  - A) Kinetic energies are increased.
  - B) Densities of the gases increased
  - C) Collisions become less frequent
  - D) Attractive force dominate at low temperature

(MCC02-UHS-08)

- An ideal gas can't be liquefied because.
  A) Its critical temperature is always above O0C
  B) Its molecules are smaller in size
  C) It solidified before becoming a liquid
  D) Forces between its molecules are negligible

(MCC02-UHS-09)

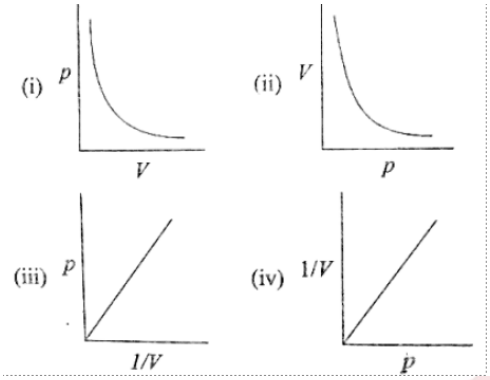
- Who attributed the deviation of real gases from ideal behavior
  A) Boyle's law
  B) Charles's law
  C) Avogadro's law
  D) Van der Waals

(MCC02-UHS-10)

- If P, V, T represent pressure, volume and temperature of the gas, the correct representation of Boyle's law is
  A) V is inversely proportional to P (at constant T)
  B) PV = nRT (at const T)
  C) Reciprocal Volume is inversely proportional to T(at constant P)
  D) The quotient of P and V is equal to K

(MCC02-UHS-11)

- Which graph represents Boyle's law?



A) (i),(ii),(iii),(iv)
B) (i),(ii),(iv)
C) (i),(iii)
D) (ii),(iii)

(MCC02-UHS-12)

- For a definite mass of an ideal gas at constant temperature, the plot of which pair of species will gives a curved graph
  A) PV vs P
  B) P vs V
  C) P vs V -1
  D) PV vs P -1

(MCC02-UHS-13)

- When sample of a gas is compressed at constant temperature from 15 atm to 60 atm, its volume changes from 76.0cm^3 to 20.5cm^3?
  A) The gas behaves ideally
  B) The gas behaves non-ideally
  C) The volume of gas decreases
  D) Gas is absorbed on the vessel walls

(MCC02-UHS-14)

- What volume does 400 cm^3 sample of a gas at 760 torr occupy when the pressure is changed to 2 atm?
  - A) 200 cm^3
  - B) 184 cm^3
  - C) 800 cm^3
  - D) 0.02 dm^3
  - E) 0.4dm^3

(MCC02-UHS-15)

- The gas is at 3 atm, what will be the pressure of a gas if it expands three times.
  - A) 1 atm
  - B) 3 atm
  - C) 6 atm
  - D) 9 atm

  ANS: (A)

Answer Key	
1	A
2	D
3	B
4	B
5	C
6	D
7	D
8	D
9	D
10	A
11	A
12	B
13	B
14	A
15	A