

(ECC02-0003H)

- Ammonia gas diffuse twice time faster than other gas. The most probable gas is:
  - A)  $\text{SO}_2$
  - B)  $\text{C}_4\text{H}_{10}$
  - C)  $\text{C}_5\text{H}_8$
  - D)  $\text{Cl}_2$

(ECC02-0004E)

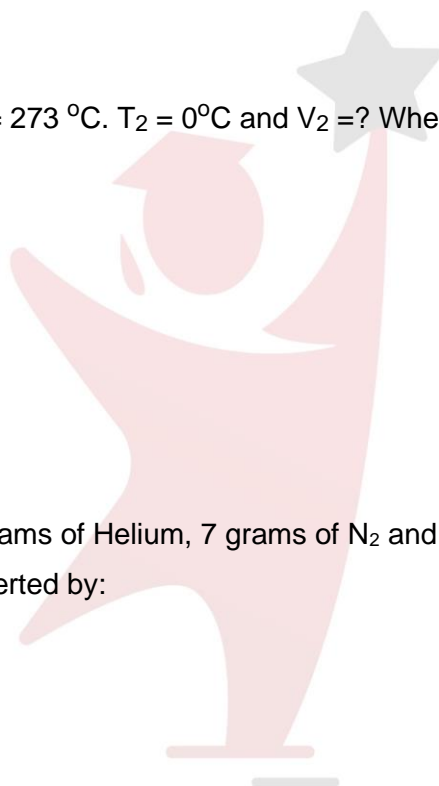
- If  $V_1 = 5$  litres,  $P_1 = 2$  atm,  $T_1 = 273$  °C.  $T_2 = 0$ °C and  $V_2 = ?$  When  $P_2 = 1$  atm.
  - A) 5 lit
  - B) 10 lit
  - C) 2.5 lit
  - D) 12.5 lit

(ECG02-0004M)

- When 1 gram of Hydrogen, 4 grams of Helium, 7 grams of  $\text{N}_2$  and 8 grams of  $\text{O}_2$  are present in a vessel, maximum partial pressure is exerted by:
  - A) Hydrogen
  - B) Helium
  - C) Nitrogen
  - D) Oxygen

(ECC02-0005E)

- K.E.  $\propto T$  is expression for
  - A) Boyle's Law
  - B) Charles's law
  - C) Kinetic Molecular Theory
  - D) None



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(ECC02-0007E)

- The spontaneous homogenous mixing of molecules of different gases by random motion and collision called
  - A) Diffusion
  - B) Effusion
  - C) Dalton's law
  - D) None

(ECC02-0008E)

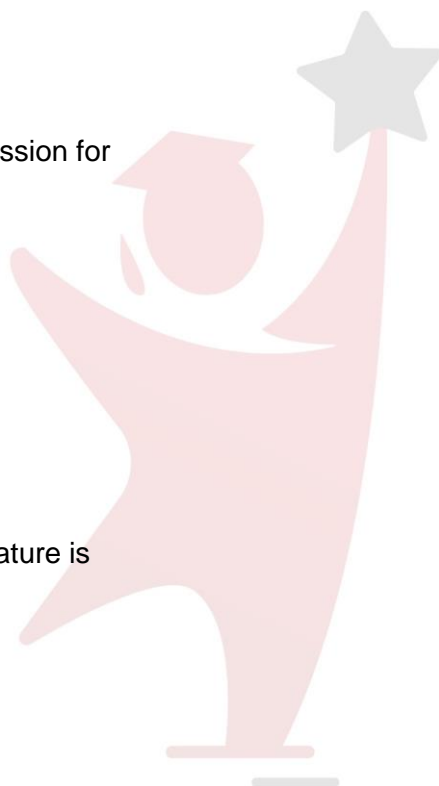
- $V \propto 1/P$ ; is mathematical expression for
  - A) Charles's Law
  - B) Avogadro's Law
  - C) Boyle's Law
  - D) Brown's Law

(ECC02-0013M)

- Most ideal gas at room temperature is
  - A)  $\text{CO}_2$
  - B)  $\text{NH}_3$
  - C)  $\text{SO}_2$
  - D)  $\text{N}_2$

(ECC02-0016E)

- Which of the following pair has same numbers of molecules at STP
  - A)  $1000\text{cm}^3$  of  $\text{N}_2\text{H}_4$  and  $\text{O}_2$
  - B)  $200\text{cm}^3$  of  $\text{CO}_2$  and  $\text{N}_2\text{O}$
  - C)  $50\text{cm}^3$  of  $\text{CO}$  and  $\text{N}_2$
  - D) All above



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(ECC02-0017M)

- The deviation of a gas from ideal behavior is maximum at
  - A)  $-10^{\circ}\text{C}$  and 5.0 atm
  - B)  $-10^{\circ}\text{C}$  and 2.0 atm
  - C)  $100^{\circ}\text{C}$  and 2.0 atm
  - D)  $0^{\circ}\text{C}$  and 2.0 atm

(ECC02-0018H)

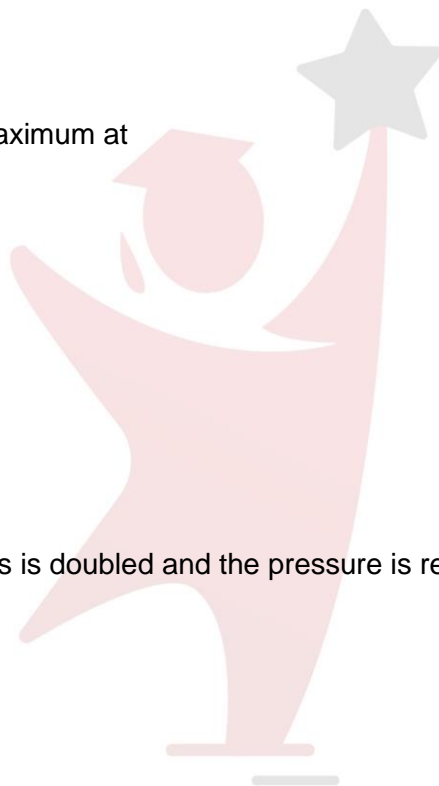
- The molar volume of  $\text{CO}_2$  is maximum at
  - A) STP
  - B)  $127^{\circ}\text{C}$  and 1 atm
  - C)  $0^{\circ}\text{C}$  and 1 atm
  - D)  $273^{\circ}\text{C}$  and 1 atm

(ECC02-0019H)

- If absolute temperature of a gas is doubled and the pressure is reduced to half the volume of gas will
  - A) Remain unchanged
  - B) Increases four times
  - C) Reduce to  $1/4^{\text{th}}$
  - D) Be doubled

(ECC02-0020E)

- The order of the rate of diffusion of gases  $\text{NH}_3$   $\text{SO}_2$   $\text{Cl}_2$  and  $\text{CO}_2$  is
  - A)  $\text{NH}_3 > \text{SO}_2 > \text{Cl}_2 > \text{CO}_2$
  - B)  $\text{NH}_3 > \text{CO}_2 > \text{SO}_2 > \text{Cl}_2$
  - C)  $\text{Cl}_2 > \text{SO}_2 > \text{CO}_2 > \text{NH}_3$
  - D) None of them



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(ECC02-0021M)

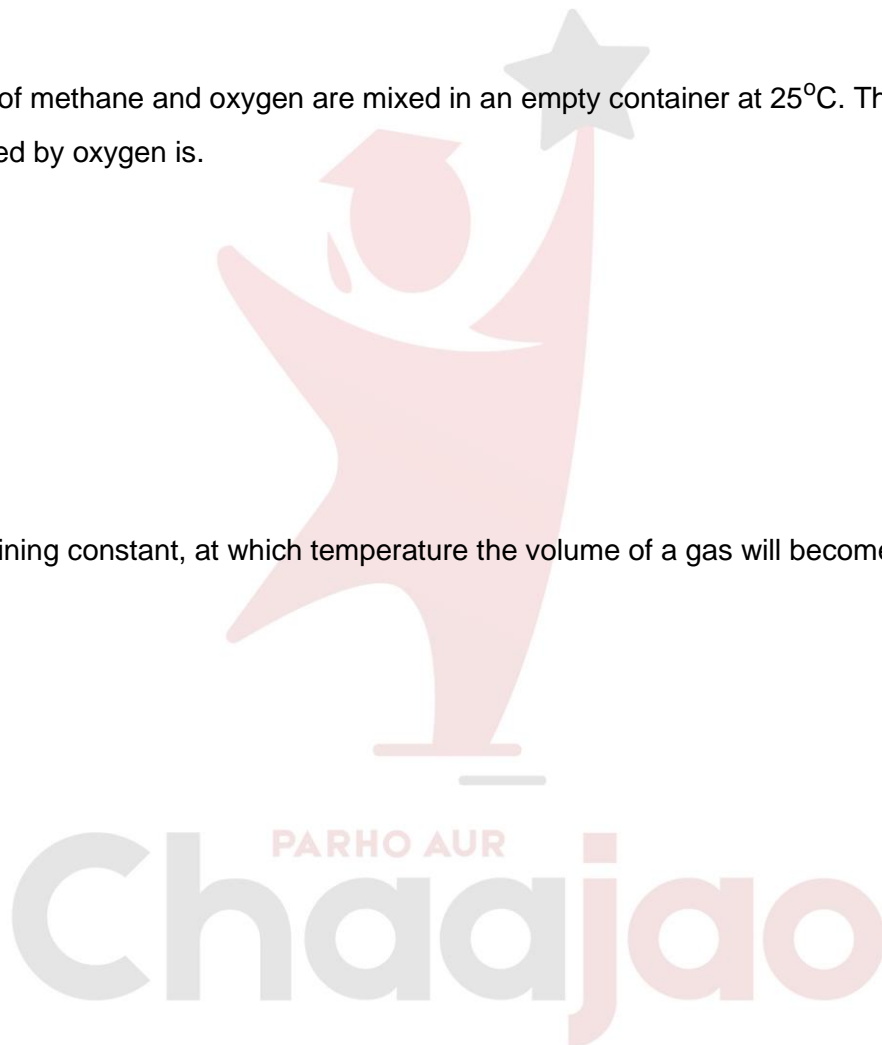
- Linde's method is employed for.
  - A) Separation of gases
  - B) Expansion of gases
  - C) Compression of gases
  - D) Liquefaction of gases

(ECC02-0022H)

- Equal masses of methane and oxygen are mixed in an empty container at 25°C. The fraction of total pressure exerted by oxygen is.
  - A) 1/3
  - B) 8/9
  - C) 1/9
  - D) 16/17

(ECC02-0024E)

- Pressure remaining constant, at which temperature the volume of a gas will become twice of what is at 0°C.
  - A) 546°C
  - B) 200°C
  - C) 546K
  - D) 273K



Answer Key	
1	C
2	A
3	B
4	C
5	A
6	C
7	D
8	D
9	A
10	D
11	B
12	B
13	D
14	A
15	C

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