(ECC02-NE-01E)

- The theory which explains that gases consist of molecules, which are in rapid motion is known as:
  A) Dalton's atomic theory
  B) Bohr's theory
  C) Rutherford's atomic theory
  D) Kinetic molecular theory

(ECC02-NE-02E)

- The law, which states that at constant temperature, the volume of a given mass of gas is inversely proportional to pressure, is known as:
  A) Boyle's law
  B) Charles' law
  C) Combine gas law
  D) Avogadro's law

(ECC02-NE-03E)

- To maintain a constant pressure, molecules with decreased average velocity will
  A) Occupy larger space
  B) Occupy smaller space
  C) Shrink
  D) None of these

ECC02-NE-04E

- Graham's law refers to:
  A) Boiling point of water
  B) Gaseous diffusion
  C) Gas compression
  D) Volume changes of gases

(ECC02-NE-06E)

- For a given mass of at constant pressure, the volume is directly proportional to absolute temperature. This is known as:
  A) Boyle's law
  B) Charles' law
  C) Avogadro's law
  D) Gay Lussac's law

(ECC02-NE-07E)

- Absolute zero refers to:
  A) 0^0 C
  B) -100^0 C
  C) -273^0 C
  D) -373^0 C

(ECC02-NE-08E)

- General gas equation is a combination of :
  A) Boyle's law, Charles' law and Avogadro's law
  B) Charles' law, Boyle's law and Dalton's law
  C) Boyle's law, Charles' law and Graham's law
  D) Charles' law, Graham's law and Avogadro's law

(ECC02-NE-09E)

- The constant temperature at which the solid and liquid phases of a substance are in equilibrium is called
  A) Freezing point
  B) Melting point
  C) Both A & B
  D) None of the above

(ECC02-NE-10E)

- The rates of diffusion of gases are inversely proportional to square root of their densities. This statement refers to
  A) Dalton's law
  B) Graham's law
  C) Avogadro's law
  D) None of the above

(ECC02-NE-11E)

- $\frac{T_1}{T2} = \sqrt{\frac{m_2}{m_1}}$ is mathematical representation of:
  A) Charles' law
  B) Combine gas law
  C) Graham's law
  D) Avogadro's law

(ECC02-NE-12E)

- Gases deviate from ideal behavior at:
  A) High temperature
  B) Low temperature and high pressure
  C) Low pressure
  D) None of them

(ECC02-NE-13E)

- Pressure has marked effect on the volume of:
  A) Gases
  B) Liquids
  C) Solids
  D) All of three states of matter equally

(ECC02-NE-14E)

- Which of the following involve strengthening of attraction between the molecules
  A) Crystallization
  B) Condensation
  C) Freezing
  D) All of these

(ECC02-NE-15E)

- The state of matter having no definite shape but definite volume
  A) Gas
  B) Liquid
  C) Solid
  D) None of the above

Answer Key	
1	**D**
2	**A**
3	**B**
4	**B**
5	**B**
6	**C**
7	**A**
8	**C**
9	**B**
10	**C**
11	**B**
12	**D**
13	**B**
14	**B**