

(ECC02-NE-01E)

- The theory which explains that gases consist of molecules, which are in rapid motion is known as:
  - A) Dalton's atomic theory
  - B) Bohr's theory
  - C) Rutherford's atomic theory
  - D) Kinetic molecular theory

(ECC02-NE-02E)

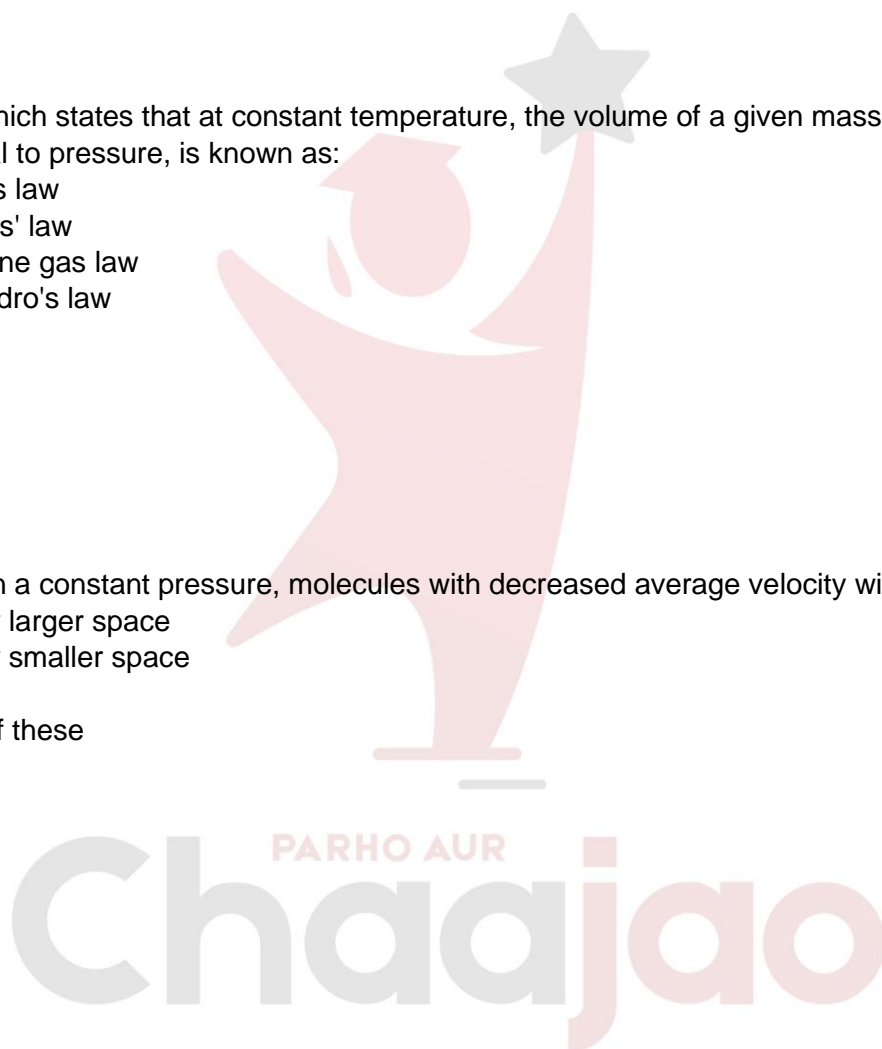
- The law, which states that at constant temperature, the volume of a given mass of gas is inversely proportional to pressure, is known as:
  - A) Boyle's law
  - B) Charles' law
  - C) Combine gas law
  - D) Avogadro's law

(ECC02-NE-03E)

- To maintain a constant pressure, molecules with decreased average velocity will
  - A) Occupy larger space
  - B) Occupy smaller space
  - C) Shrink
  - D) None of these

ECC02-NE-04E

- Graham's law refers to:
  - A) Boiling point of water
  - B) Gaseous diffusion
  - C) Gas compression
  - D) Volume changes of gases



(ECC02-NE-06E)

- For a given mass of at constant pressure, the volume is directly proportional to absolute temperature. This is known as:  
A) Boyle's law  
B) Charles' law  
C) Avogadro's law  
D) Gay Lussac's law

(ECC02-NE-07E)

- Absolute zero refers to:  
A)  $0^{\circ}\text{C}$   
B)  $-100^{\circ}\text{C}$   
C)  $-273^{\circ}\text{C}$   
D)  $-373^{\circ}\text{C}$

(ECC02-NE-08E)

- General gas equation is a combination of :  
A) Boyle's law, Charles' law and Avogadro's law  
B) Charles' law, Boyle's law and Dalton's law  
C) Boyle's law, Charles' law and Graham's law  
D) Charles' law, Graham's law and Avogadro's law

(ECC02-NE-09E)

- The constant temperature at which the solid and liquid phases of a substance are in equilibrium is called  
A) Freezing point  
B) Melting point  
C) Both A & B  
D) None of the above



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(ECC02-NE-10E)

- The rates of diffusion of gases are inversely proportional to square root of their densities. This statement refers to
  - A) Dalton's law
  - B) Graham's law
  - C) Avogadro's law
  - D) None of the above

(ECC02-NE-11E)

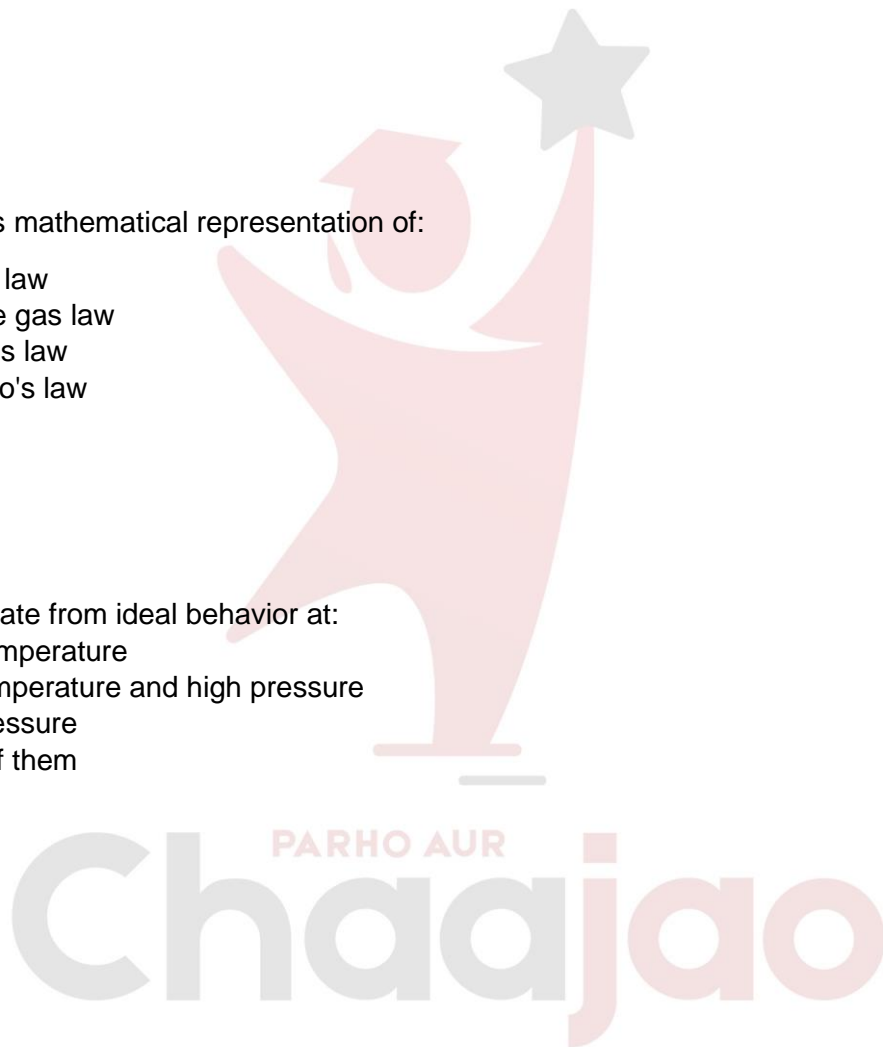
- $\frac{T_1}{T_2} = \sqrt{\frac{m_2}{m_1}}$  is mathematical representation of:
  - A) Charles' law
  - B) Combine gas law
  - C) Graham's law
  - D) Avogadro's law

(ECC02-NE-12E)

- Gases deviate from ideal behavior at:
  - A) High temperature
  - B) Low temperature and high pressure
  - C) Low pressure
  - D) None of them

(ECC02-NE-13E)

- Pressure has marked effect on the volume of:
  - A) Gases
  - B) Liquids
  - C) Solids
  - D) All of three states of matter equally

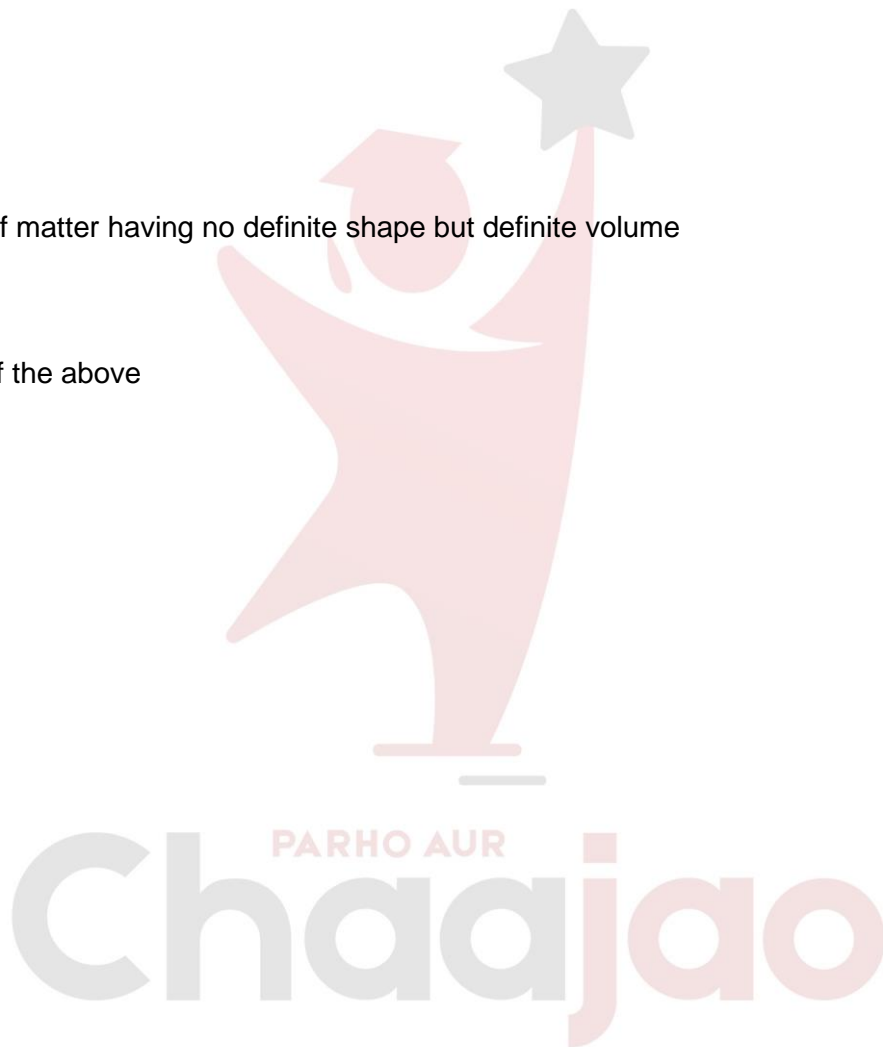


(ECC02-NE-14E)

- Which of the following involve strengthening of attraction between the molecules
  - A) Crystallization
  - B) Condensation
  - C) Freezing
  - D) All of these

(ECC02-NE-15E)

- The state of matter having no definite shape but definite volume
  - A) Gas
  - B) Liquid
  - C) Solid
  - D) None of the above



Answer Key	
1	D
2	A
3	B
4	B
5	B
6	C
7	A
8	C
9	B
10	C
11	B
12	D
13	B
14	B

Chaaajao