

(ECC01-0001H)

- The positive ions having group of atoms are _____ common.
A) More
B) Less
C) Equal
D) None

(ECC01-0002E)

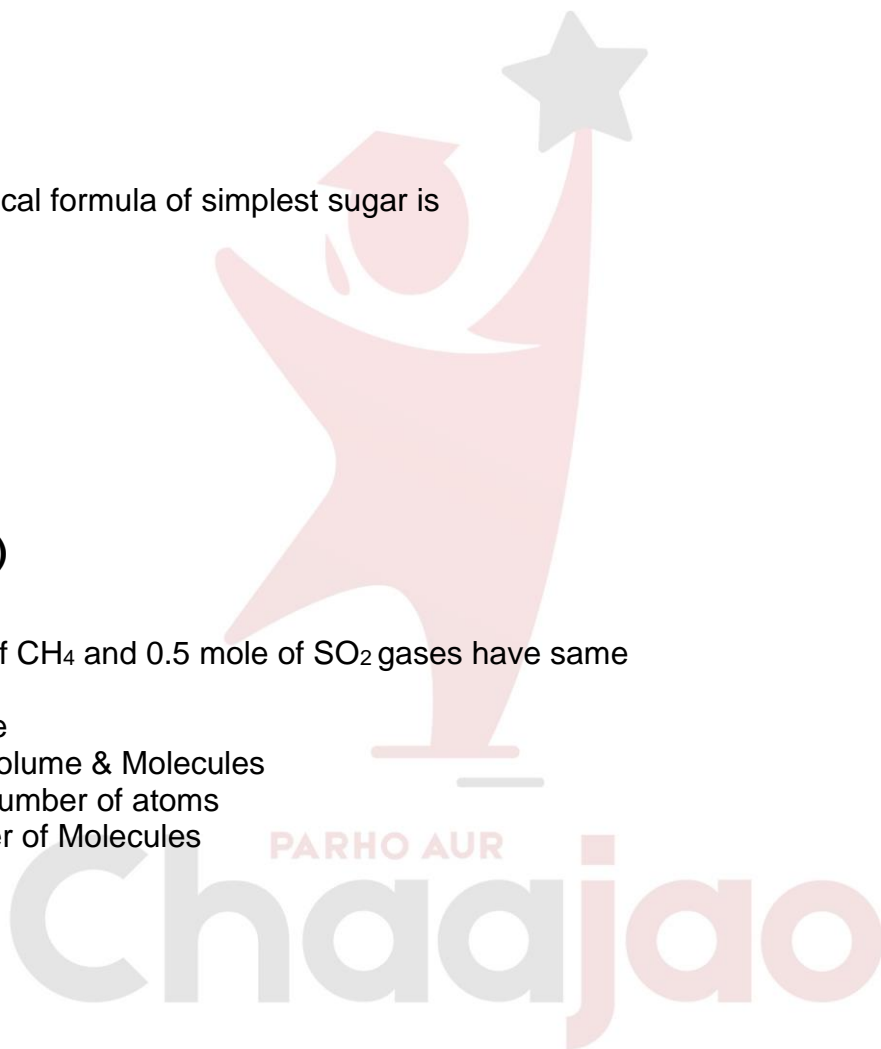
- The empirical formula of simplest sugar is
a. CH
b. CH₂O
c. C₆H₆
d. CHO

(ECC01-0003M)

- 0.5 mole of CH₄ and 0.5 mole of SO₂ gases have same
A) Volume
B) Both Volume & Molecules
C) Total number of atoms
D) Number of Molecules

(ECC01-0004M)

- What is the mass of water formed when 4 grams H₂ and 64 grams of O₂ combined together
A) 66 grams
B) b. 18 grams
C) 36 grams
D) d. 66 grams



(ECC01-0005H)

- The total number of covalent bonds in 4.5 grams of water is
 - A) 6.02×10^{23}
 - B) 6.02×10^{22}
 - C) 3.01×10^{22}
 - D) 3.01×10^{23}

(ECC01-0006E)

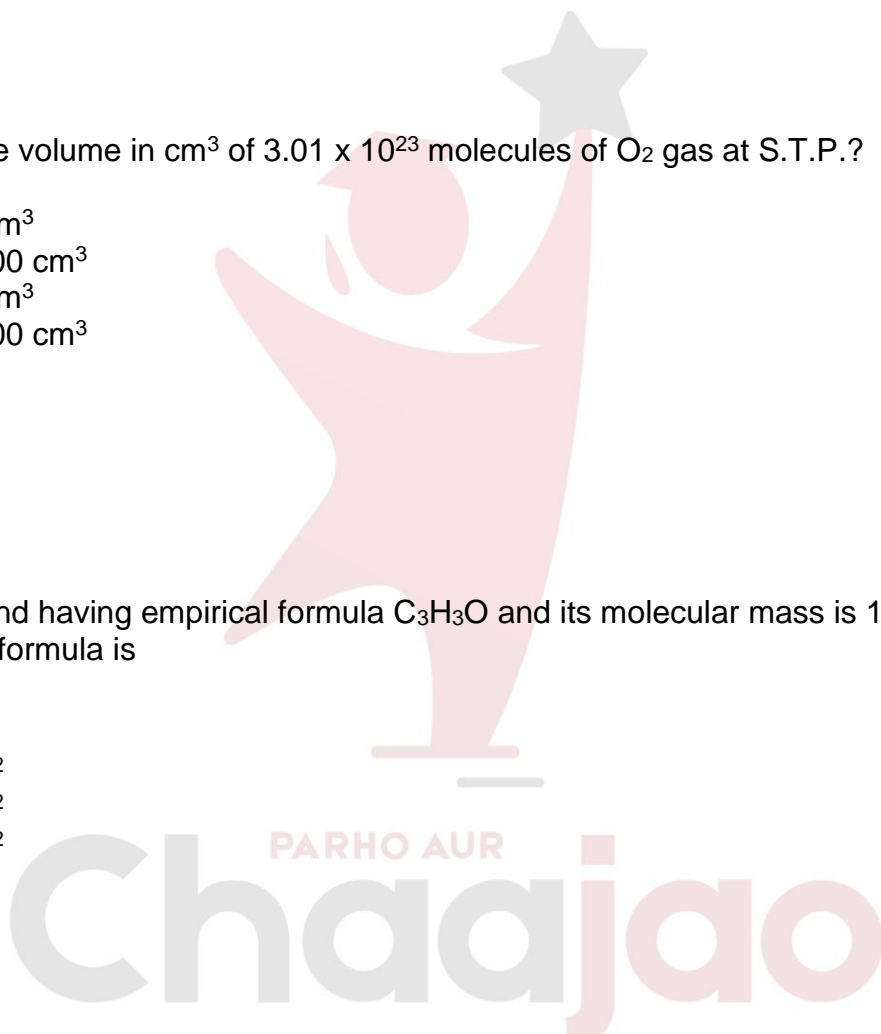
- What is the volume in cm^3 of 3.01×10^{23} molecules of O_2 gas at S.T.P.?
 - A) 1000 cm^3
 - B) 11000 cm^3
 - C) 1120 cm^3
 - D) 11200 cm^3

(ECC01-0007E)

- A compound having empirical formula $\text{C}_3\text{H}_3\text{O}$ and its molecular mass is 110.02 its molecular formula is
 - A) $\text{C}_3\text{H}_3\text{O}$
 - B) $\text{C}_6\text{H}_6\text{O}_2$
 - C) $\text{C}_9\text{H}_9\text{O}_2$
 - D) $\text{C}_3\text{H}_6\text{O}_2$

(ECC01-0008M)

- In the determination of atomic ratio of the elements the mole ratios are divided by
 - A) Least value of gram atoms of elements
 - B) Atomic masses of elements
 - C) Given mass of the Compound
 - D) Molecular mass of the compound



(ECC01-0009E)

- Which one of the following steps is not involved in determination of empirical formula?
 - A) Determination % of each element
 - B) Determination of gram atoms of each elements
 - C) Determination of isotopes of each elements
 - D) Determination of atomic ratio of elements

(ECC01-0010M)

- Which of the following salts will have greater positive charge in its molar aqueous solution keeping in mind that solution as a whole remains neutral?
 - A) CaCl_2
 - B) KCl
 - C) NH_4Cl
 - D) NaCl

(ECC01-0011E)

- A molecule is the smallest particle of a substance because
 - A) It has positive charge on it.
 - B) It exists independently
 - C) It decomposes into ions
 - D) It is always mono atomic

(ECC01-0012M)

- 9.8 grams of aqueous solution of H_2SO_4 contains moles of H^+ ions
 - A) 0.1
 - B) 0.2
 - C) 0.11
 - D) 0.01

(ECC01-0013M)

- The pressure of vapours when sent to the ionization chamber in mass spectrometer is
 - A) 10^{-5} to 10^{-6} torr
 - B) 10^{-6} to 10^{-7} torr
 - C) 10^{-7} to 10^{-8} torr
 - D) 10^{-3} to 10^{-4} torr

(ECC01-0014E)

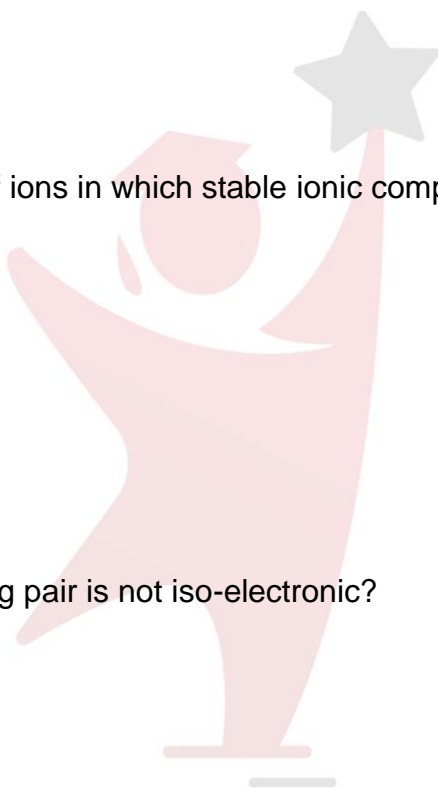
- The smallest collection of ions in which stable ionic compound is stable:
 - A) Chemical formula
 - B) Molecular formula
 - C) Formula unit
 - D) Formula mass

(ECC01-0015E)

- Which one of the following pair is not iso-electronic?
 - A) CO, N₂
 - B) Na⁺, Ne
 - C) Ca, Ar
 - D) K⁺, Ar

(ECC01-0016H)

- Which one of the following is not a molecular ion?
 - A) N₂⁺
 - B) CH₄⁺
 - C) C₇H₈⁺
 - D) NH₄⁺



PARHO AUR
Chaaajao

(ECC01-0017H)

- Atoms of the elements must be containing in nucleus
 - A) Proton
 - B) Proton and neutron
 - C) Neutron
 - D) Electron and neutron

(ECC01-0018M)

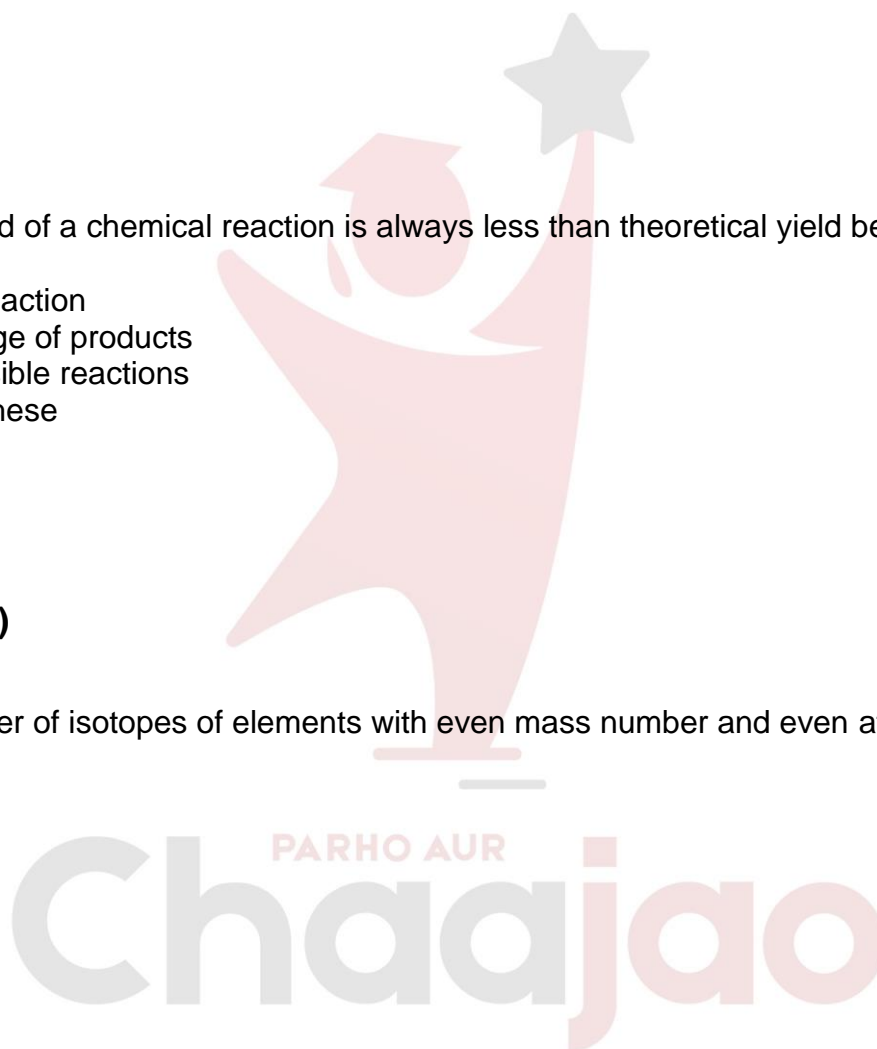
- Actual yield of a chemical reaction is always less than theoretical yield because
 - A) Side reaction
 - B) Wastage of products
 - C) Reversible reactions
 - D) All of these

(ECC01-0019M)

- The number of isotopes of elements with even mass number and even atomic number are
 - A) 280
 - B) 300
 - C) 154
 - D) 54

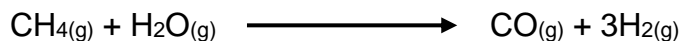
(ECC01-0020M)

- Which one of the following is not the mono isotopic elements?
 - A) Arsenic
 - B) Uranium
 - C) Iodine
 - D) Gold



(ECC01-0021M)

- Methane reacts with steam to form CO and H₂



What volume of hydrogen gas can be obtained from 100 cm³ of methane at STP?

- A) 300 cm³
- B) 200 cm³
- C) 150 cm³
- D) 100 cm³

(ECC01-0022E)

- 1.4 gm of N₂ reacts with hydrogen, what will be volume of NH₃ at S.T.P?

- A) 224 cm³
- B) 22.4 dm³
- C) 2.24 dm³
- D) 1.12 dm³

(ECC01-0023M)

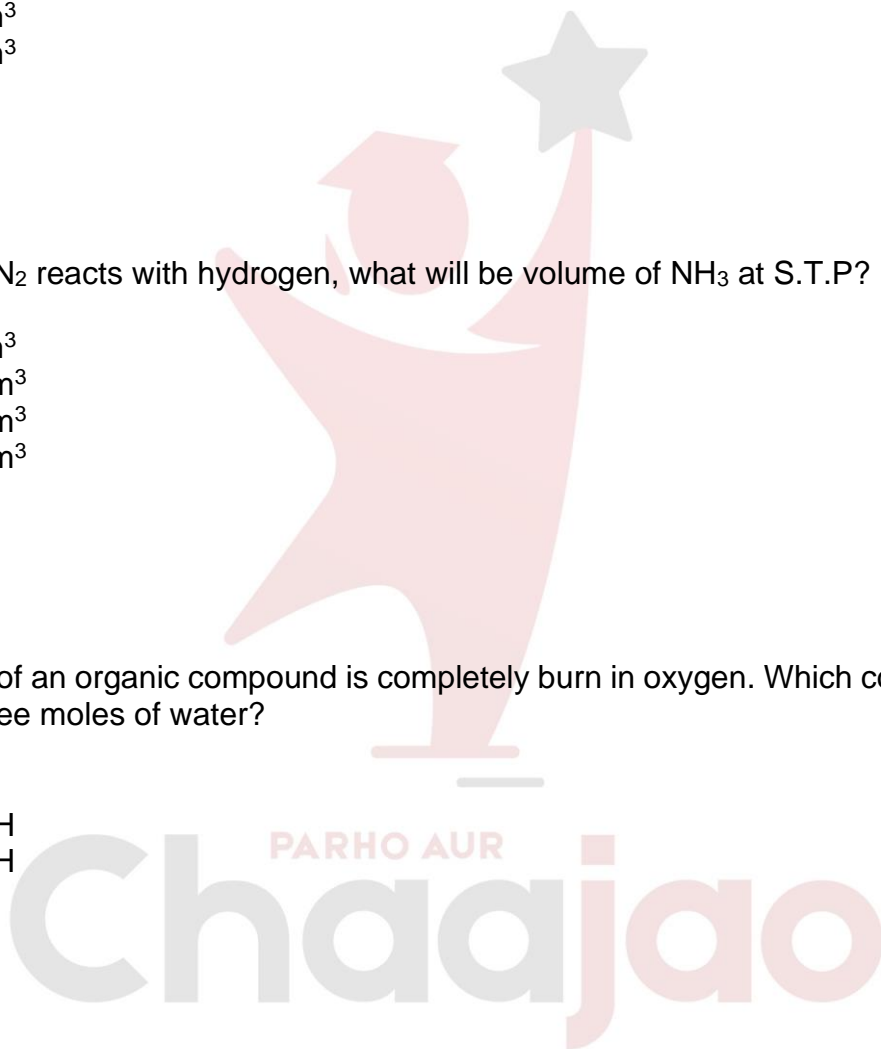
- One mole of an organic compound is completely burn in oxygen. Which compound produces exactly three moles of water?

- A) C₄H₁₀
- B) C₄H₉OH
- C) C₂H₅OH
- D) C₃H₈

(ECC01-0024M)

- Which of the following contains the same number of molecules as 9g of water?

- A) 2g of hydrogen gas
- B) 14g of nitrogen gas
- C) 32g of oxygen gas
- D) 44g of carbon dioxide



(ECC01-0025H)

- An element E forms a hydride EH_3 which contains 90.0% of E, by mass. What is the relative atomic mass of E?

- A) 27
- B) 87
- C) 30
- D) 90



Answers Key	
1	B
2	B
3	D
4	C
5	D
6	D
7	B
8	A
9	C
10	A
11	B
12	B
13	B
14	C
15	C
16	D
17	A
18	D
19	C
20	B
21	A
22	C
23	C
24	B
25	A

ChaaJao



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