(ECC01-NE-01E)

- One mole of C2H4 reacts with three moles of O2. 56g of C2H4 will react with
  A) 6 moles of O2
  B) 3.61 x 1024 molecules of O2
  C) 134.4 dm3 of O2
  D) All of these

(ECC01-NE-02E)

- 5894 is written to base 10 as:
  A) 5.894 x 10-3
  B) 5.894 x 104
  C) 5.894 x 102
  D) 5.894 x 103

(ECC01-NE-03E)

- 870.0 have _____ significant figures.
  A) Two
  B) Three
  C) Four
  D) Six

(ECC01-NE-04E)

- A chemical formula based on actual number of molecule is called _____ .
  A) Structural
  B) Molecular
  C) Empirical
  D) None

(ECC01-NE-05E)

- Which of the following is a single pure compound?
  A) Air
  B) Steam
  C) Brass
  D) Sea Water

(ECC01-NE-06E)

- Formula of ferric sulphate is:
  A) FeSO4
  B) Fe(SO4)3
  C) Fe2(SO4)3
  D) Fe2SO4

(ECC01-NE-07E)

- A mixture whose composition is uniform throughout the mixture is called:
  A) Heterogeneous
  B) Homogenous
  C) Homonuclear
  D) Heteronuclear

(ECC01-NE-08E)

- The number of moles present in 6gms of carbon is:
  A) 2
  B) 0.5
  C) 5
  D) 1

(ECC01-NE-09E)

- 1 cm3 is equal to _____ m3:
  A) 10^{-2}
  B) 10^{-6}
  C) 10^{6}
  D) 10^{4}

(ECC01-NE-10E)

- _____16 gm of O2 contains
  A) 3.01 X 10^{23} molecules
  B) 6.02 X 10^{23} atoms
  C) 6.02 X 10^{24} molecules
  D) 3.01 X 10^{-23} atoms

(ECC01-NE-11E)

- The -ve charged particles is called:
  A) Anion
  B) Cation
  C) Radical
  D) Atom

(ECC01-NE-12E)

- 20g of a monoatomic gas occupies 44.8 dm3 at STP. Its atomic weight is
  A) 5g
  B) 10g
  C) 40g
  D) 20g

(EPC01-NE-13E)

- a pure chemical compound always contains the same elements combined in the same ratio by weight is:
  A) Law of definite proportions
  B) Law of multiple proportions
  C) Law of mass action
  D) Law of equilibrium

(ECC01-NE-14E)

- Any charged particle is called:
  A) Atom
  B) Molecule
  C) Ion
  D) Mixture

(ECC01-NE-15E)

- The no. of moles present in 1200 cm^3 of a gas at S.T.P are:
  A) 0.52 moles
  B) 1.51 moles.
  C) 0.053 moles
  D) 0.053 moles

Answer Key	
1	**D**
2	**D**
3	**C**
4	**B**
5	**B**
6	**C**
7	**B**
8	**B**
9	**B**
10	**A**
11	**A**
12	**B**
13	**A**
14	**C**
15	**C**