

(ECC01-NE-01E)

- One mole of  $C_2H_4$  reacts with three moles of  $O_2$ . 56g of  $C_2H_4$  will react with
  - A) 6 moles of  $O_2$
  - B)  $3.61 \times 10^{24}$  molecules of  $O_2$
  - C) 134.4 dm<sup>3</sup> of  $O_2$
  - D) All of these

(ECC01-NE-02E)

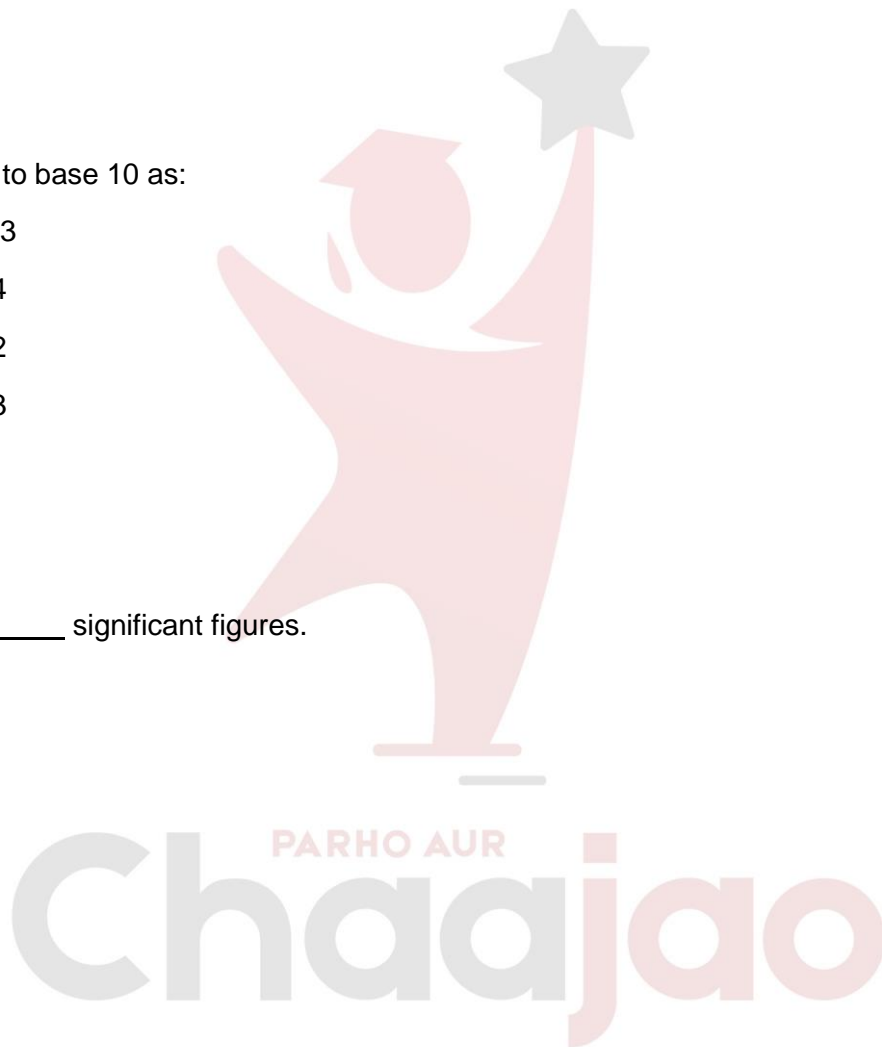
- 5894 is written to base 10 as:
  - A)  $5.894 \times 10^{-3}$
  - B)  $5.894 \times 10^4$
  - C)  $5.894 \times 10^2$
  - D)  $5.894 \times 10^3$

(ECC01-NE-03E)

- 870.0 have \_\_\_\_\_ significant figures.
  - A) Two
  - B) Three
  - C) Four
  - D) Six

(ECC01-NE-04E)

- A chemical formula based on actual number of molecule is called \_\_\_\_\_.
  - A) Structural
  - B) Molecular
  - C) Empirical
  - D) None



(ECC01-NE-05E)

- Which of the following is a single pure compound?
  - A) Air
  - B) Steam
  - C) Brass
  - D) Sea Water

(ECC01-NE-06E)

- Formula of ferric sulphate is:
  - A)  $\text{FeSO}_4$
  - B)  $\text{Fe}(\text{SO}_4)_3$
  - C)  $\text{Fe}_2(\text{SO}_4)_3$
  - D)  $\text{Fe}_2\text{SO}_4$

(ECC01-NE-07E)

- A mixture whose composition is uniform throughout the mixture is called:
  - A) Heterogeneous
  - B) Homogenous
  - C) Homonuclear
  - D) Heteronuclear

(ECC01-NE-08E)

- The number of moles present in 6gms of carbon is:
  - A) 2
  - B) 0.5
  - C) 5
  - D) 1



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(ECC01-NE-09E)

- 1 cm<sup>3</sup> is equal to \_\_\_\_\_ m<sup>3</sup>:  
A) 10<sup>-2</sup>  
B) 10<sup>-6</sup>  
C) 10<sup>6</sup>  
D) 10<sup>4</sup>

(ECC01-NE-10E)

- \_\_\_\_\_ 16 gm of O<sub>2</sub> contains  
A) 3.01 X 10<sup>23</sup> molecules  
B) 6.02 X 10<sup>23</sup> atoms  
C) 6.02 X 10<sup>24</sup> molecules  
D) 3.01 X 10<sup>-23</sup> atoms

(ECC01-NE-11E)

- The -ve charged particles is called:  
A) Anion  
B) Cation  
C) Radical  
D) Atom

(ECC01-NE-12E)

- 20g of a monoatomic gas occupies 44.8 dm<sup>3</sup> at STP. Its atomic weight is  
A) 5g  
B) 10g  
C) 40g  
D) 20g



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(EPC01-NE-13E)

- a pure chemical compound always contains the same elements combined in the same ratio by weight is:
  - A) Law of definite proportions
  - B) Law of multiple proportions
  - C) Law of mass action
  - D) Law of equilibrium

(ECC01-NE-14E)

- Any charged particle is called:
  - A) Atom
  - B) Molecule
  - C) Ion
  - D) Mixture

(ECC01-NE-15E)

- The no. of moles present in  $1200 \text{ cm}^3$  of a gas at S.T.P are:
  - A) 0.52 moles
  - B) 1.51 moles.
  - C) 0.053 moles
  - D) 0.053 moles



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| Answer key |   |
|------------|---|
| 1          | D |
| 2          | D |
| 3          | C |
| 4          | B |
| 5          | B |
| 6          | C |
| 7          | B |
| 8          | B |
| 9          | B |
| 10         | A |
| 11         | A |
| 12         | B |
| 13         | A |
| 14         | C |
| 15         | C |

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