

(ECC01-NE-01E)

- One mole of C_2H_4 reacts with three moles of O_2 . 56g of C_2H_4 will react with
 - A) 6 moles of O_2
 - B) 3.61×10^{24} molecules of O_2
 - C) 134.4 dm^3 of O_2
 - D) All of these

(ECC01-NE-02E)

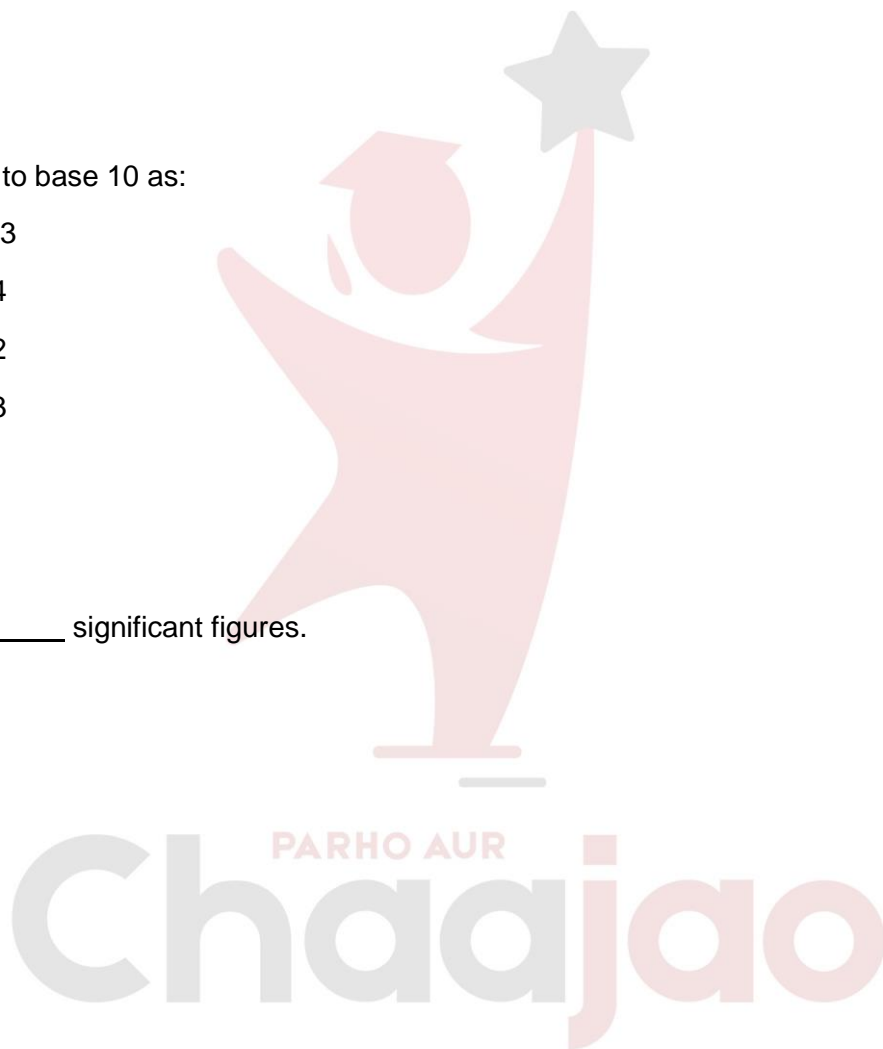
- 5894 is written to base 10 as:
 - A) 5.894×10^{-3}
 - B) 5.894×10^4
 - C) 5.894×10^2
 - D) 5.894×10^3

(ECC01-NE-03E)

- 870.0 have _____ significant figures.
 - A) Two
 - B) Three
 - C) Four
 - D) Six

(ECC01-NE-04E)

- A chemical formula based on actual number of molecule is called _____.
 - A) Structural
 - B) Molecular
 - C) Empirical
 - D) None



(ECC01-NE-05E)

- Which of the following is a single pure compound?
 - A) Air
 - B) Steam
 - C) Brass
 - D) Sea Water

(ECC01-NE-06E)

- Formula of ferric sulphate is:
 - A) FeSO_4
 - B) $\text{Fe}(\text{SO}_4)_3$
 - C) $\text{Fe}_2(\text{SO}_4)_3$
 - D) Fe_2SO_4

(ECC01-NE-07E)

- A mixture whose composition is uniform throughout the mixture is called:
 - A) Heterogeneous
 - B) Homogenous
 - C) Homonuclear
 - D) Heteronuclear

(ECC01-NE-08E)

- The number of moles present in 6gms of carbon is:
 - A) 2
 - B) 0.5
 - C) 5
 - D) 1



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(ECC01-NE-09E)

- 1 cm³ is equal to _____ m³:
A) 10⁻²
B) 10⁻⁶
C) 10⁶
D) 10⁴

(ECC01-NE-10E)

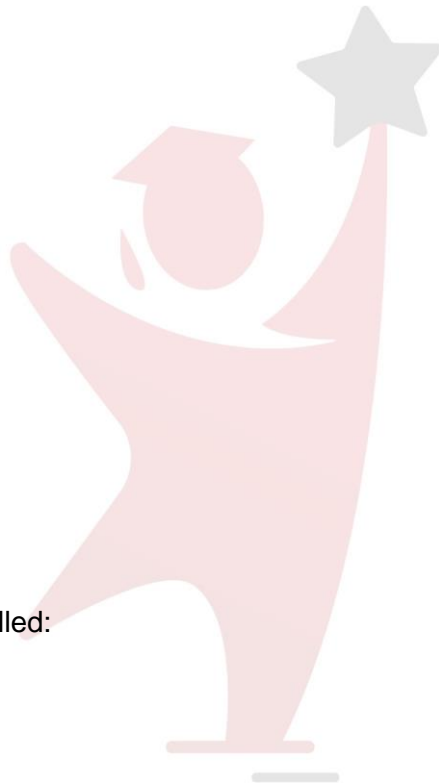
- _____ 16 gm of O₂ contains
A) 3.01 X 10²³ molecules
B) 6.02 X 10²³ atoms
C) 6.02 X 10²⁴ molecules
D) 3.01 X 10⁻²³ atoms

(ECC01-NE-11E)

- The -ve charged particles is called:
A) Anion
B) Cation
C) Radical
D) Atom

(ECC01-NE-12E)

- 20g of a monoatomic gas occupies 44.8 dm³ at STP. Its atomic weight is
A) 5g
B) 10g
C) 40g
D) 20g



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(EPC01-NE-13E)

- a pure chemical compound always contains the same elements combined in the same ratio by weight is:
 - A) Law of definite proportions
 - B) Law of multiple proportions
 - C) Law of mass action
 - D) Law of equilibrium

(ECC01-NE-14E)

- Any charged particle is called:
 - A) Atom
 - B) Molecule
 - C) Ion
 - D) Mixture

(ECC01-NE-15E)

- The no. of moles present in 1200 cm^3 of a gas at S.T.P are:
 - A) 0.52 moles
 - B) 1.51 moles.
 - C) 0.053 moles
 - D) 0.053 moles



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Answer key	
1	D
2	D
3	C
4	B
5	B
6	C
7	B
8	B
9	B
10	A
11	A
12	B
13	A
14	C
15	C

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