(ECC01-0002E)

- The empirical formula of simplest sugar is

  A) CH
  B) CH_2O
  C) C_6H_6
  D) CHO

(ECC01-0003M)

- 0.5 mole of CH_4 and 0.5 mole of SO_2 gases have same

  A) Volume
  B) Both Volume & Molecules
  C) Total number of atoms
  D) Number of Molecules

(ECC01-0004M)

- What is the mass of water formed when 4 grams H_2 and 64 grams of O_2 combined together

  A) 66 grams
  B) 18 grams
  C) 36 grams
  D) 66 grams
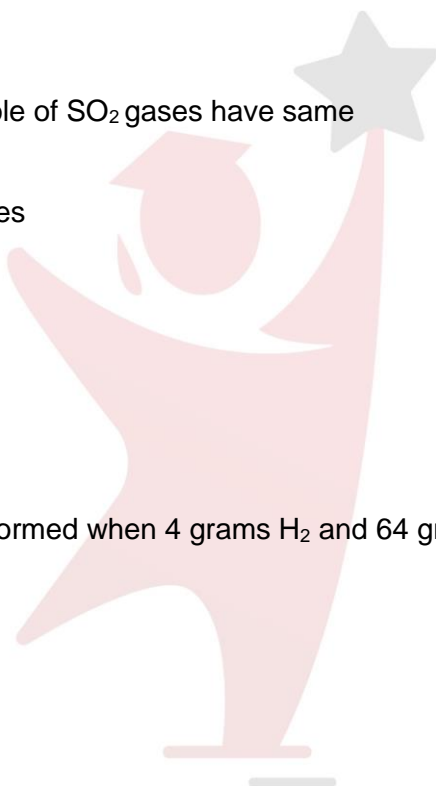
(ECC01-0005H)

- The total number of covalent bonds in 4.5 grams of water is

  A) 6.02×10^{23}
  B) 6.02×10^{22}
  C) 3.01×10^{22}
  D) 3.01×10^{23}

(ECC01-0006E)

- What is the volume in cm^3 of 3.01×10^{23} molecules of O_2 gas at S.T.P.?

  A) $1000 \ cm^3$
  B) $11000 \ cm^3$
  C) $1120 \ cm^3$
  D) $11200 \ cm^3$

(ECC01-0007E)

- A compound having empirical formula C_3H_3O and its molecular mass is 110.02 its molecular formula is

  A) C_3H_3O
  B) $C_6H_6O_2$
  C) $C_9H_9O_2$
  D) $C_3H_6O_2$

(ECC01-0008M)

- In the determination of atomic ratio of the elements the mole ratios are divided by

  A) Least value of gram atoms of elements
  B) Atomic masses of elements
  C) Given mass of the Compound
  D) Molecular mass of the compound

(ECC01-0009E)

- Which one of the following steps is not involved in determination of empirical formula?

  A) Determination % of each element
  B) Determination of gram atoms of each elements
  C) Determination of isotopes of each elements
  D) Determination of atomic ratio of elements

(ECC01-0011E)

- A molecule is the smallest particle of a substance because

  A) It has positive charge on it.
  B) It exists independently
  C) It decomposes into ions
  D) It is always mono atomic

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(ECC01-0012M)

- 9.8 grams of aqueous solution of H_2SO_4 contains moles of H^+ ions

 A) 0.1
  B) 0.2
  C) 0.11
  D) 0.01

(ECC01-0014E)

- The smallest collection of ions in which stable ionic compound is stable:

  A) Chemical formula
  B) Molecular formula
  C) Formula unit
  D) Formula mass

(ECC01-0015E)

- Which one of the following pair is not iso-electronic?

  A) CO, N_2
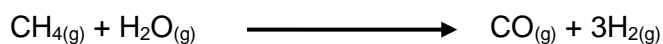 B) Na^+, Ne
  C) Ca, Ar
  D) K^+, Ar

(ECC01-0018M)

- Actual yield of a chemical reaction is always less than theoretical yield because

  A) Side reaction
  B) Wastage of products
  C) Reversible reactions
  D) All of these

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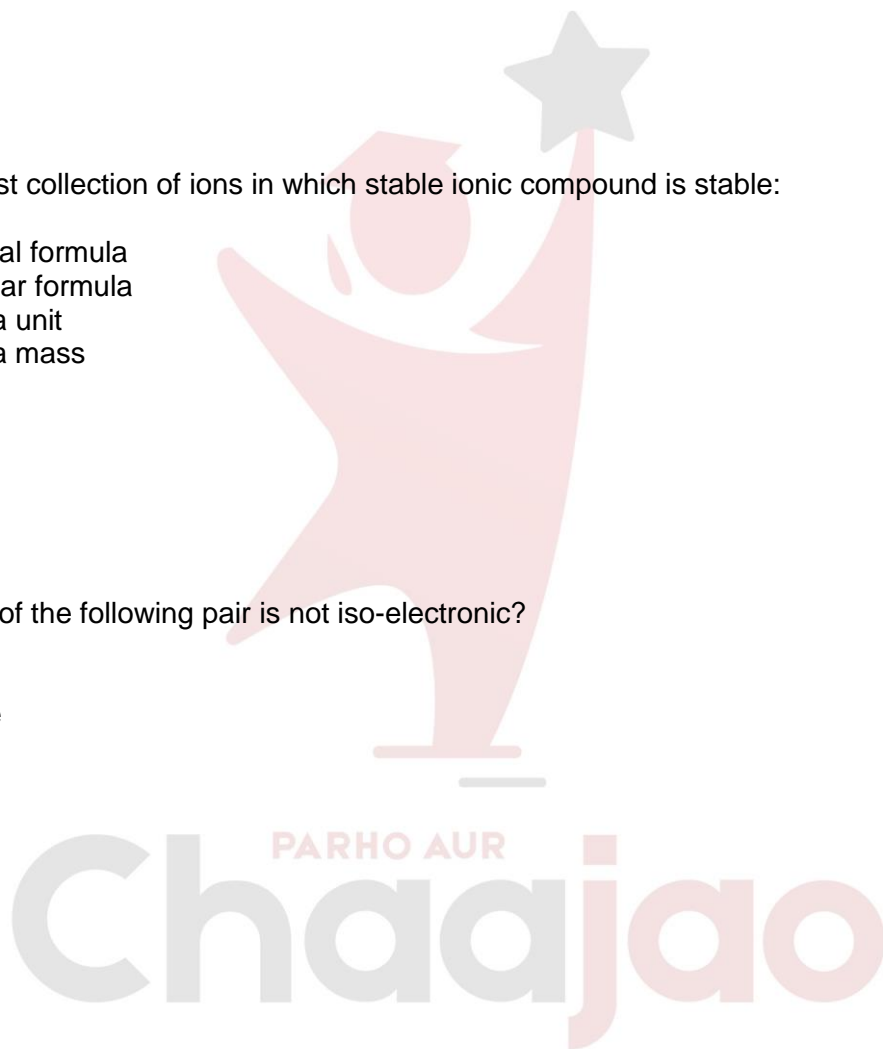
(ECC01-0021M)

- Methane reacts with steam to form CO and H_2

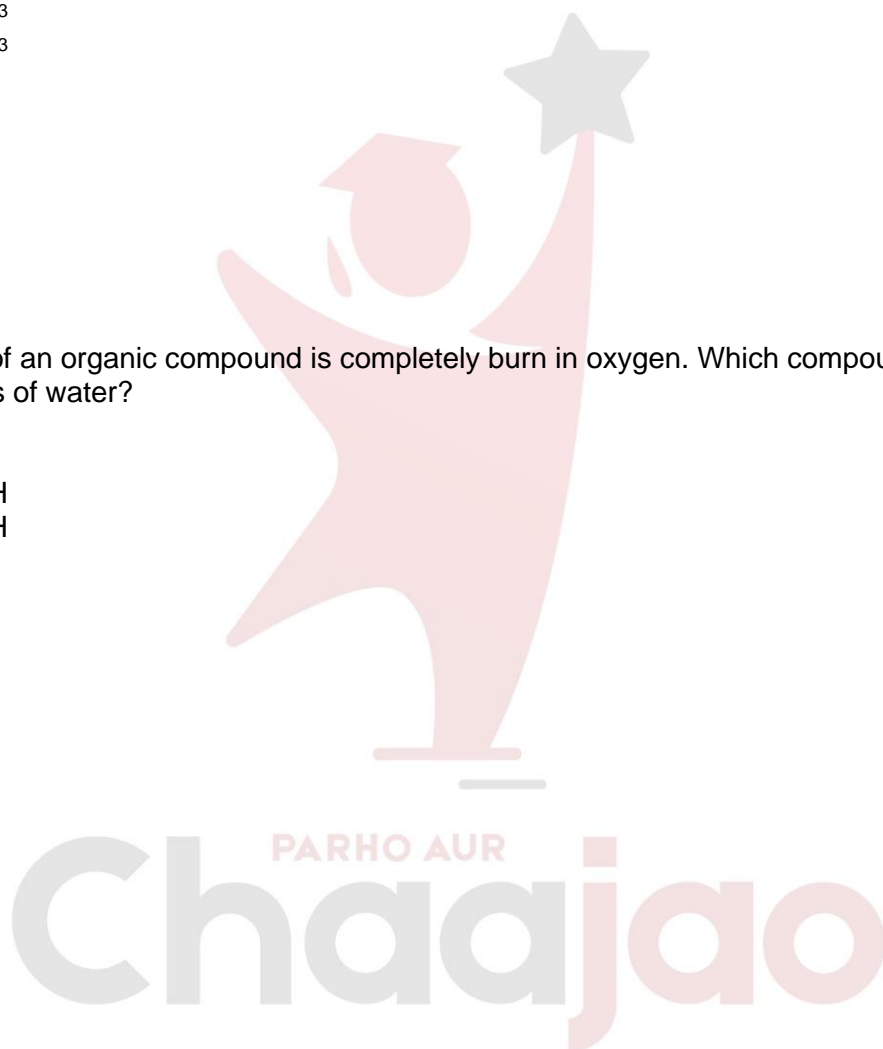$$CH_{4(g)} + H_2O_{(g)} \longrightarrow CO_{(g)} + 3H_{2(g)}$$

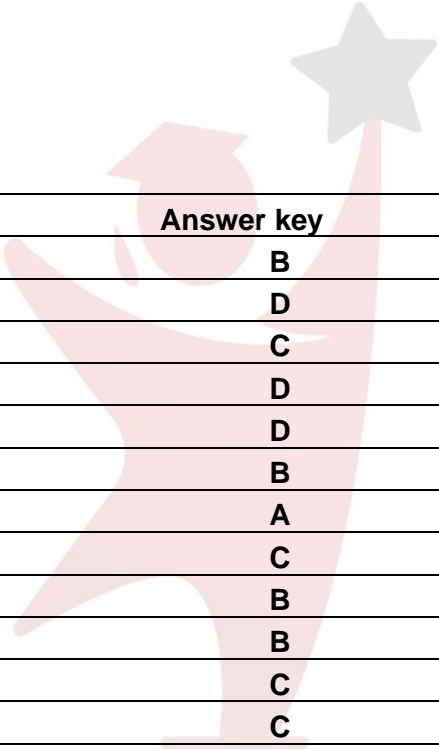What volume of hydrogen gas can be obtained from 100 cm^3 of methane at STP?

A) 300 cm^3
B) 200 cm^3
C) 150 cm^3
D) 100 cm^3

(ECC01-0023M)

- One mole of an organic compound is completely burn in oxygen. Which compound produces exactly three moles of water?

A) C_4H_{10}
B) C_4H_9OH
C) C_2H_5OH
D) C_3H_8

Answer key		
1		B
2		D
3		C
4		D
5		D
6		B
7		A
8		C
9		B
10		B
11		C
12		C
13		D
14		A
15		C

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