

(ECC01-0001H)

- The positive ions having group of atoms are _____ common.
A) More
B) Less
C) Equal
D) none

(ECC01-0002E)

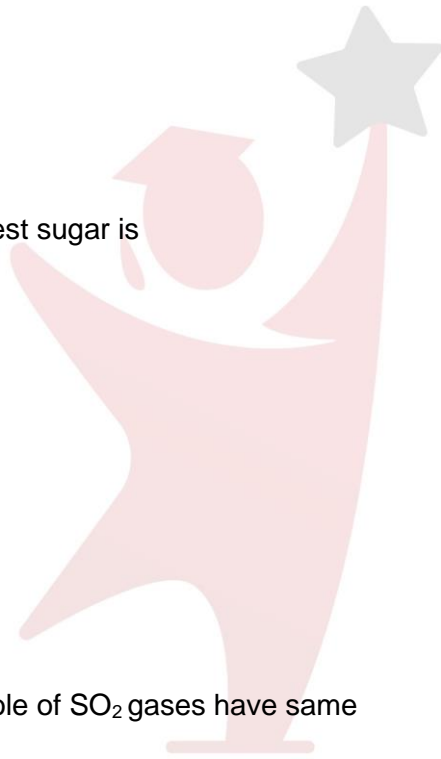
- The empirical formula of simplest sugar is
A) CH
B) CH₂O
C) C₆H₆
D) CHO

(ECC01-0003M)

- 0.5 mole of CH₄ and 0.5 mole of SO₂ gases have same
A) Volume
B) Both Volume & Molecules
C) Total number of atoms
D) Number of Molecules

(ECC01-0004M)

- What is the mass of water formed when 4 grams H₂ and 64 grams of O₂ combined together
A) 66 grams
B) 18 grams
C) 36 grams
D) 66 grams



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(ECC01-0005H)

- The total number of covalent bonds in 4.5 grams of water is

- A) 6.02×10^{23}
- B) 6.02×10^{22}
- C) 3.01×10^{22}
- D) 3.01×10^{23}

(ECC01-0006E)

- What is the volume in cm^3 of 3.01×10^{23} molecules of O_2 gas at S.T.P.?

- A) 1000 cm^3
- B) 11000 cm^3
- C) 1120 cm^3
- D) 11200 cm^3

(ECC01-0002E-PMC-01M)

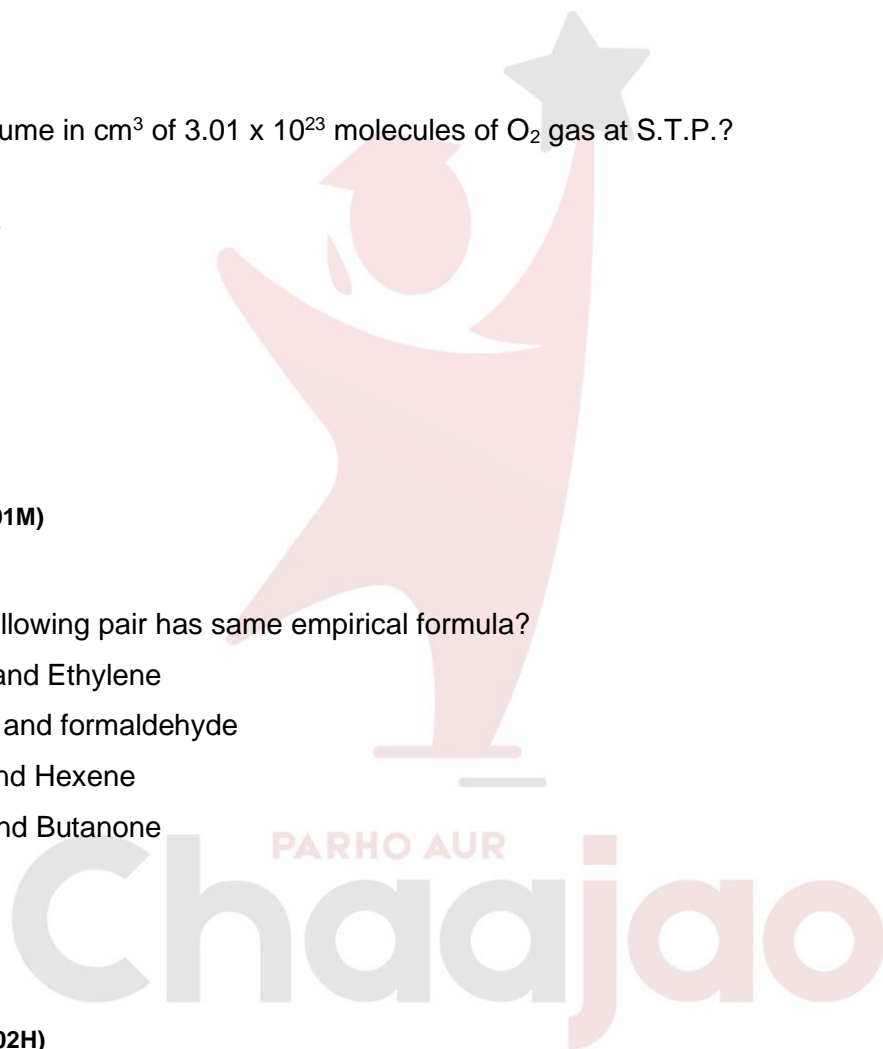
- Which of the following pair has same empirical formula?

- A) Acetylene and Ethylene
- B) Acetic acid and formaldehyde
- C) Benzene and Hexene
- D) Propanal and Butanone

(ECC01-0003M-PMC-02H)

- 0.5 mole of H_2O and 0.5 mole of SO_2 gases have same.

- A) Molecules
- B) Both volume & Molecules
- C) Atoms, volume and molecules
- D) Both Atoms & Molecules



(ECC01-0004M-PMC-03H)

- If 49 grams of H_2SO_4 react with 80.0 grams of NaOH , how much reactant will be left over after the reaction is complete?
 - 24.5 g H_2SO_4
 - none of either compound
 20. g NaOH
 40. g NaOH

(ECC01-0006E-PMC-04E)

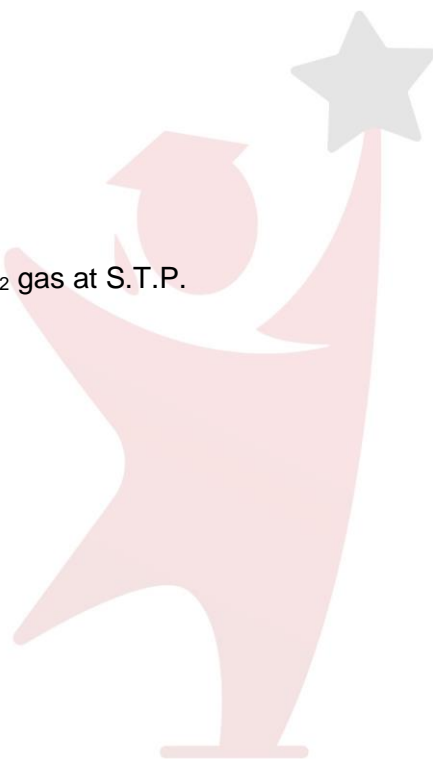
- the molecules in 5600 cm^3 of O_2 gas at S.T.P.
 - 3.01×10^{23}
 - 3.01×10^{22}
 - 6.01×10^{22}
 - 1.5×10^{23}

(ECC01-0012M-PMC-05H)

- 9.8 grams of aqueous solution of H_3PO_4 contains total moles of ions
 - 0.1
 - 0.2
 - 0.3
 - 0.4

(ECC01-0013M-PMC-06M)

- How many moles of neutron are present in one mole of heavy water?
 - 18
 - 8
 - 10
 - 20



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(ECC01-0019M-PMC-07H)

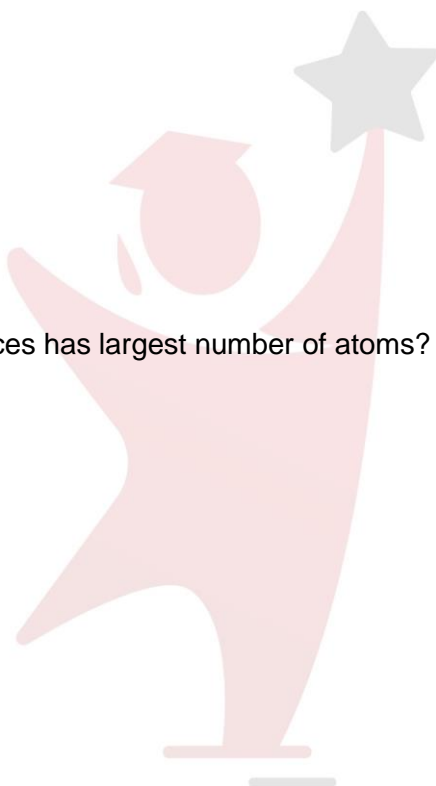
- The anion in calcium hydrogen phosphate
 - A) $[\text{H}_3\text{PO}_4]^{-1}$
 - B) $[\text{H}_2\text{PO}_4]^{-1}$
 - C) $[\text{PO}_4]^{-3}$
 - D) $[\text{HPO}_4]^{-2}$

(ECC01-0021M-PMC-08M)

- Which of the following substances has largest number of atoms?
 - A) 7g sulphur dioxide gas
 - B) 7g magnesium metal
 - C) 7g oxygen gas
 - D) 7g carbon dioxide gas

(ECC01-0022E-PMC-09E)

- 0.8 gm of CH_4 reacts with oxygen, what will be volume of CO_2 at S.T.P?
 - A) 224 cm^3
 - B) 22.4 dm^3
 - C) 2.24 dm^3
 - D) 1.12 dm^3



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Answer key	
1	B
2	B
3	D
4	C
5	D
6	D
7	B
8	D
9	D
10	D
11	D
12	C
13	D
14	D
15	D

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