

(ECC02-0001E)

- The collisions between gas molecules are
 - A) Elastic
 - B) Inelastic
 - C) Both a and b
 - D) None

(ECC02-0002H)

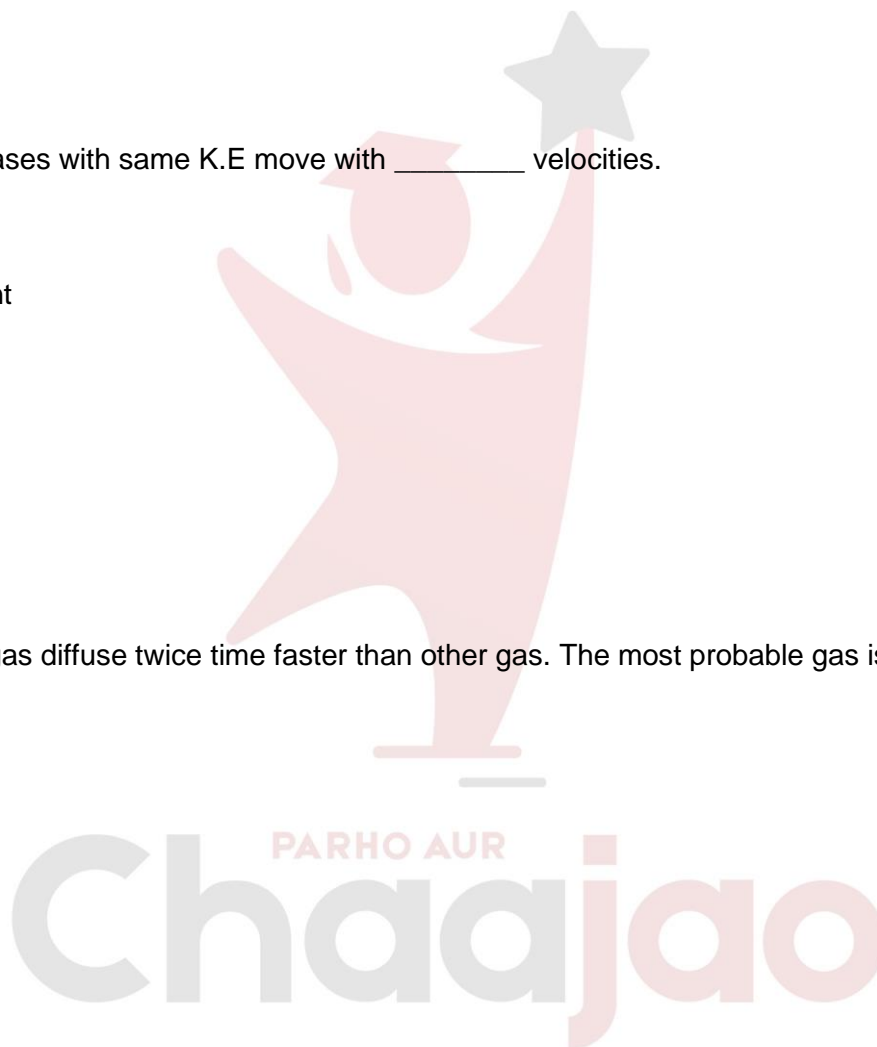
- Different gases with same K.E move with _____ velocities.
 - A) Same
 - B) Different
 - C) Equal
 - D) None

(ECC02-0003H)

- Ammonia gas diffuse twice time faster than other gas. The most probable gas is:
 - A) SO₂
 - B) C₄H₁₀
 - C) C₅H₈
 - D) Cl₂

(ECC02-0004E)

- If $V_1 = 5$ litres, $P_1 = 2$ atm, $T_1 = 273$ °C. $T_2 = 0$ °C and $V_2 = ?$ When $P_2 = 1$ atm.
 - A) 5 lit
 - B) 10 lit
 - C) 2.5 lit
 - D) 12.5 lit



(ECC02-0005E)

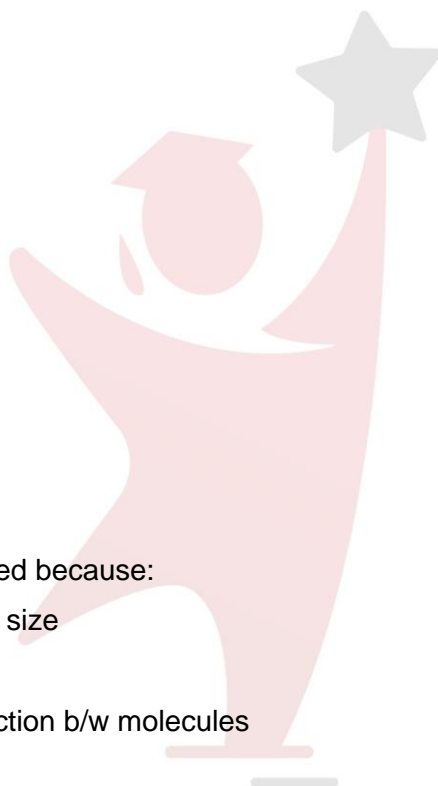
- K.E. $\propto T$ is expression for
 - A) Boyle's Law
 - B) Charles's law
 - C) Kinetic Molecular Theory
 - D) None

(ECC02-0006M)

- Absolute zero is equal to
 - A) 0°C
 - B) -459.7°F
 - C) -273°K
 - D) none

(ECC02-0016E-PMC-01E)

- An ideal gas cannot be liquified because:
 - A) Its molecules are small in size
 - B) Show no mass
 - C) There is no force of attraction b/w molecules
 - D) Both A & B



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(ECC02-0007E-PMC-02M)

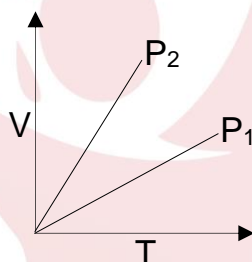
- If the four tubes of a car are filled to the same pressure with N_2 , O_2 , H_2 and He separately, then which one will be filled first?
 - A) N_2
 - B) O_2
 - C) H_2
 - D) He

(ECC02-0001E-PMC-03E)

- The kinetic theory of gases predicts that total kinetic energy of a gaseous assembly depends on:
 - Pressure of the gas
 - Temperature of the gas
 - Volume of the gas
 - Pressure, temperature and volume of the gas

(ECC02-0028E-PMC-04H)

- "V" versus "T" straight lines at constant pressures P_1 and P_2 for an ideal gas are shown in figure. Which is correct?



- $P_1 > P_2$
- $P_1 < P_2$
- $P_1 = P_2$
- Both B & C

(ECC02-0026M-PMC-07M)

- An ideal gas obeying kinetic theory of gases can be liquified if:
 - Its temperature is more than critical temperature
 - It can not be liquified at any P &
 - Its pressure is more than critical pressure
 - Both A & C

(ECC02-0027H-PMC-08M)

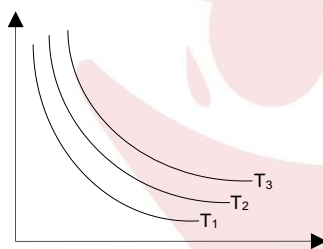
- Rate of diffusion of gas X is twice the gas Y. If molecular weight of X is 8, what is molecular weight of Y?
 - 32
 - 16
 - 8
 - 18

(ECC02-0029M-PMC-10M)

- A gas can be liquified by pressure alone when its temperature is:
 - Higher than the critical temperature
 - Lower than the critical temperature
 - Equal to its critical temperature
 - None

(ECC02-0030M-PMC-11H)

- “P” versus “V” curves and constant temperatures T_1 , T_2 , & T_3 are given in figure. Which is correct?



- $T_1 < T_2 < T_3$
- $T_1 = T_2 = T_3$
- $T_2 < T_1 < T_3$
- $T_1 > T_2 > T_3$

(ECC02-0036M-PMC-12E)

- At constant volume for a fixed number of moles of a gas, the pressure of a gas increases with the rise in temperature due to:
 - Increase in molecular attraction
 - Increase in mean free path
 - Increase in average molecular speed
 - Increase in rate of collisions amongst molecular

Answer Key	
1	A
2	B
3	C
4	A
5	C
6	B
7	C
8	C
9	B
10	A
11	B
12	A
13	B
14	A
15	C

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